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CHAPTER 8 COVALENT BONDING

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Chapter 8 Covalent Bonding Single

Covalent Bonds When only one pair of

electrons is shared, such as in a hydrogen

molecule, it is a single covalent bond. The

shared electron pair is often referred to as

the bonding pair. For a hydrogen molecule,

shown in Figure 8.4, each covalently bonded

atom equally, 1 Chapter 8: Covalent Bonding

Section 8.1 Section 8.2 Section 8.3 Section

8.4 Section 8.5 The Covalent Bond Naming

Molecules Molecular Structures Molecular

Shapes Electronegativity and Polarity

Review Vocabulary chemical bond: the force

that holds two atoms together oxyanion: a

polyatomic ion in which an element (usually a

nonmetal) is bonded to one or more oxygen

atoms ionic bond: the ..., Chapter 8 Concepts

of Chemical Bonding . Chemical Bonds

Three types: Ionic Electrostatic attraction

between ions Covalent Sharing of electrons

Metallic Metal atoms bonded to several other

atoms . Ionic Bonding ... these covalent

bonds., Chapter 8: Covalent Bonding Notes

Chemical Bonds o Atoms bond to become

more stable o Lower energy states are more

stable than higher energy states o Energy is

released when a bond forms Covalent Bonds

o Sharing of valence electrons o Valence

atomic orbitals overlap end-to-end s-s orbital

overlap, s-p orbital overlap, or p-p orbital

overlap ..., Chapter 8 Covalent Bonding and

Molecular Structure 8-1 Chapter 8: Covalent

Bonding and Molecular Structure Chapter In

Context In this chapter and the next, we

examine chemical bonding in detail., Chapter

8 Advanced Theories of Covalent Bonding

Figure 8.1 Oxygen molecules orient

randomly most of the time, as shown in the

top magnified view. However, when we pour

liquid oxygen through a magnet, the

molecules line up with the magnetic field,

and the attraction allows them, Chapter 8

Bonding: General Concepts 8 - 2 unequal

attraction results in a polar covalent bond.

The stronger atom has an overabundance of

electron, Chapter 8 Covalent bonding A

metal and a nonmetal transfer electrons

An ionic bond ... For example

show how water is formed with covalent

bonds. Each hydrogen has 1 valence electron and wants 1 more ...

Notes, Whiteboard, Whiteboard Page, Notebook software, Notebook, PDF, SMART, SMART Technologies ULC, SMART Board Interactive Whiteboard, 8 Covalent Bonding Section 8.1 The Covalent Bond In your textbook, read about the nature of covalent bonds. ... Study Guide - Chapter 8 " Covalent Bonding Section 8.1 The Covalent Bond 1. covalent bond 2. exothermic 3. molecule 4. sigma bond 5. pi bond 6. b 7. d 8. c 9. a Section 8.2 Naming Molecules, 2 4 Chemical Bonds " Chemical bond " A force that holds atoms together in a molecule or compound " Two types of chemical bonds " Ionic Bonds " Covalent Bonds 8 - Figure 8.2 5 Ionic Bond " A bond created by electrostatic attraction between oppositely, Chapter 8: Chemical Bonds (+ VSEPR) 5 Covalent Bonds Share and Share Alike 6 Covalent Bonds and Molecules " Covalent bonds form when two or more nonmetals share their electrons. The electrons are at their lowest potential energy when they are between the two nuclei

that are being joined., CHAPTER 9

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When resonance structures can be drawn, it is usually due to a multiple bond that can be in different positions. This is the case for NO_3^- . Experiment tells us that the three N-O bonds ..., COVALENT BONDING CHAPTER 7 Covalent Bonding We have seen that ionic bonding requires low ionization energies, high electron affinity and high lattice energies in ionic compounds. Properties of ionic compounds Solids at room temperature - mostly high melting (melting breaks up the crystal lattice that the

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