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COVALENT BONDING CHAPTER 8

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Chapter 2 1 Chapter 2: Atomic Structure and

Chemical Bonding © Materials

†'Molecules †'Atoms © Atoms = protons (p) + neutrons (n) + electrons (e), Halogen

bonding is a type of non-covalent interaction which does not involve the formation nor breaking of actual bonds, but rather is similar to the dipole-dipole interaction known as

hydrogen bonding., ©2011 American

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BONDING AND CHEMICAL STRUCTURE

hydrogen atom). This is the octet rule for covalent bonding, and it will prove to be

extremely important for understanding

chemical reactivity., 32 CHAPTER 1 ©

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STRUCTURE As shown in the diagram, the

3s and 3p electrons are the valence electrons of sulfur; the 3s and 3p orbitals are

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Introduction H 1s 1 F 2s 2 2p 5 H F eight outer electrons two shared six non-bonding

H F † + † Polar Covalent Bond 1.3 Bonding

in Carbon Compounds The property of carbon that makes it unique is its ability to

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we know if it worked? The reactant is, 3
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15, Polymer Structures University
Tennessee, Dept. of Materials Science and
Engineering 5 Many other organic groups
can be involved in, 1 Hybridization and
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(VB) - a molecule arises from interaction of
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by Mark Bishop, Unit 8: Chemical Bonding.

Why and how do atoms combine to form

compounds? In this unit, we will draw Lewis
structures to describe bonding and discuss
several models to describe chemical
bonding., Amines! Amines are organic
compounds containing a nitrogen
functionality! Depending upon the number of
alkyl, or aryl, groups attached to nitrogen !,
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